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answer. 157.5 g 106.5 g 51 g Yes.

Mass of reactants equals mass of products. Section 3.2 Assessment page 79 10. Classify each

example as a physical change or a chemical change. a. crushing an aluminum can physical b.

recycling used aluminum cans to make new aluminum cans

physical c. aluminum combining with oxygen to form aluminum oxide chemical 11.

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Questions 10–13 refer to the

following. (A) $+Ag + Br^- \rightarrow AgBr$

(B) $6\ ^{14}_6C \rightarrow 7\ ^{14}_7N + -1\ ^0_0e$ (C) $^{234}_{92}U \rightarrow 4\ ^{90}_{230}Th + 2\ He$ (D) $^{30}_{15}P \rightarrow 14\ Si + 1\ ^0_0e$ (E) $2HgO \rightarrow 2Hg + O_2$

10. Represents the decomposition of a compound

into its ...

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the electromagnetic spectrum
shown below. $c / (3.00 \times 10^8 \text{ m/s} \dots$

Chapter Assessment

mass of oxygen 536.41 214.58
24.36 517.47 g 40% calcium, 12%
carbon, 48% oxygen. 14.58/36.41
540.04% calcium 4.36/36.41
512.0% carbon 17.47/36.41
547.98% oxygen. Copyright ©
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51 g Yes. Mass of reactants
equals mass of products. Section
3.2 Assessment page 79 10.
Classify each example as

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